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Chemanalyse

your application to do chemistry laboratory practises

Chemanalyse

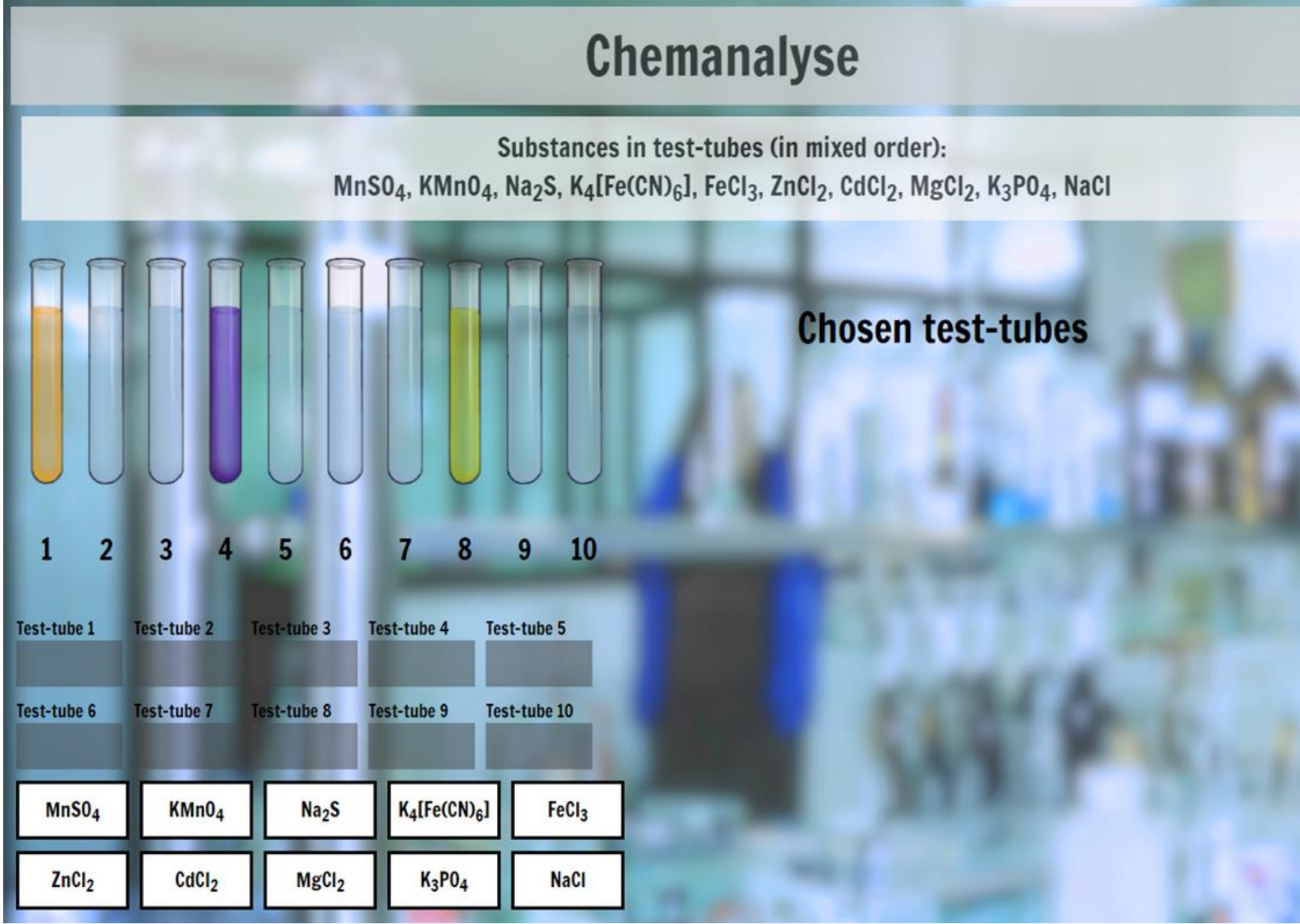
Chemistry is our joint passion. We spend a lot of time on solving difficult tasks, learning from well-known chemistry books and performing laboratory exercises and experiments. We know, that our comfortable situation of having a well-equipped laboratory within reach is rather a very seldom exception in Polish school. Even on the universities students do not have enough good access to laboratories due to lack of time and, at first, money. **We had an idea:** offering help for everybody, who wants to broaden horizons in learning chemistry, but do not have proper equipment. Our preposition is called **Chemanalyse** and it is the application one of the kind!

Mn²⁺ + MnO₄⁻ As a result of reaction Mn ²⁺ and MnO ₄ ⁻ ions does not arise manganese (II) permanganate for sure. It takes place redox and visible precipitate is manganese (IV) oxide MnO ₂ .		MnS As a result of reaction arises flesh-coloured precipitate of manganese (II) sulphide, soluble in diluted mineral acids and in CH ₃ COOH. On the air it oxidise easily to manganese (III) hydroxide and because of that colour of precipitate change into fuscous.	
Mn₂[Fe(CN)₆] As a result of reaction arises precipitate, which colour is hard to define. It is white, beige or sometimes greenish, especially when we use concentrated solutions. Because of many different colours of products reaction is not distinctive.		Mn₃(PO₄)₂ As a result of reaction arises white precipitate of manganese (II) phosphate, soluble in mineral acids and also in acetic acid.	
Fe₂S₃ In test-tube arises black precipitate of iron (III) sulphide. It is necessary to precipitate it in neutral or alkaline solution, because this precipitate is easily soluble in wide range of acids. During dissolving occurs reduction Fe ³⁺ to Fe ²⁺ and oxidation S ²⁻ to colloidal sulphur, what can we see as opacification and dark yellow precipitate.		ZnS In test-tube arises white precipitate of zinc sulphide, which arises in form of sol. It is worth remembering, that precipitates of sulphides are usually black.	

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Substances in test-tubes (in mixed order):
MnSO₄, KMnO₄, Na₂S, K₄[Fe(CN)₆], FeCl₃, ZnCl₂, CdCl₂, MgCl₂, K₃PO₄, NaCl

Chosen test-tubes



MnSO ₄	KMnO ₄	Na ₂ S	K ₄ [Fe(CN) ₆]	FeCl ₃
ZnCl ₂	CdCl ₂	MgCl ₂	K ₃ PO ₄	NaCl

Our application is dedicated not only for people, who already had caught the bug and gain some knowledge in chemistry. Thanks to **extra materials**, which we prepared in addition to application, you can start your story from basic principles to more complicated tasks. We use in application two different types of exercises – with **six or ten test-tubes**, which are filled in solutions of salts, acids or bases. In future it will be also **indicators and organic compounds**. There are tasks with different levels of challenge, so the choice is quite wide. You can also use a mood **“blind tries”**, in which you can show **all reactions**, which are available in Chemanalyse – also that reactions, which are not presented in tasks.

Chemanalyse can be used by almost everyone who is interested in chemistry. The application could be an useful educational tool for students who do not have a chance to learn chemistry in practice. Chemanalyse can become a well substitute!